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Titration

Empirical

Formula of

Magnesium Oxide

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Conclusion --

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~~PRACTICAL TEST:~~

~~LAB REPORT~~ *Molar*

Mass of Butane

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Experiment #13:

The Equilibrium

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~~[SK015] Exp 2:
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ion~~

~~Determination of
The~~

~~Concentration of
HCl Solution~~

~~(Week 3 \u0026~~

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Experiment 2 :

Acid-base
titration

Finding and

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~~CHEM 131 L
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Answers to Post-
Lab Questions
for Experiment
5. I have 40 L
of methane gas
measured at
standard
temperature and

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pressure (STP) .

How many moles
of methane gas
is this? 1 mole
of any gas at
STP = 22.4 L

Moles of methane
= $(40 / 22.4) =$
1.79 moles. What
is the volume of
0.30 moles of
nitrogen gas if
measured at STP?
1 mole of any

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gas at STP =
22.4 L Volume of
nitrogen = $0.3 \times$
 $22.4 = 6.72$ L.

~~SCC1123~~
~~Experiment 5~~
~~Post-Lab~~
~~Questions~~
~~Answers - ECU~~

~~---~~

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(Chemistry of

Copper) Part 1:

[Copper metal to

Copper (II)

nitrate) 2pts.

Balanced

chemical

equation:

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Observations:

(type your observations below after watching the video titled Formation of copper (1) nitrate. Part 2: [Preparation of Copper (11) hydroxide] 2pts.

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~~(Chemistry Of
Copper) Part 1~~

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First calculate
the moles of KHP
(potassium
hydrogen
phthalate,

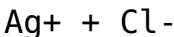
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KHC₈H₄O₄). The
molar mass for
KHP = 204.2
g/mole. 0.531 g
KHP x (1 mole
KHP / 204.2 g
KHP) = 0.00260
moles KHP.
KHP...

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The white precipitate most likely came from the presence of chloride ions.



-->AgCl(s) The chemical responsible for the iron is Fe (iron), in some form. It could be almost...

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series of
questions. When
you can answer "
yes: " to a
question, then

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stop. 1) Does your reaction have oxygen as one of it's reactants and carbon dioxide and water as products? Yes: it's a combustion reaction 2) Does your reaction have two (or more) chemicals

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combining to
form one
chemical? Yes:
it's a synthesis
reaction

~~Chemistry of
Copper~~

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a series of acid-
base
extractions.

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from the
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evaluation of
experimental
data.

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