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mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP), one mole of any gas occupies approximately 22.4 liters.

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CHEMISTRY I (TESC 141) STUDY GUIDE MOLES/ STOICHIOMETRY

Mole-a unit of measurement that expresses the amount of atoms, molecules or some other unit. The number of items in one mole is commonly referred to Avogadro's

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number which equals 6.022×10^{23} .

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Avogadro's principle tells us that 1 mole is equal to 6.022×10^{23} atoms of a pure substance. This will be important to know when converting

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from grams to moles to atoms. Now that we have learned the basics, let's try converting a sample of 50.0 grams of CO₂ between units. CO₂ -
Converting from Grams to Moles

**Moles and Molar Mass | Unit 1 -
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The answers are listed below. 1. 4.00 moles (6.02×10^{23} molecules / 1 mole) = 2.41×10^{24} molecules. 2. 0.750 moles Zn (6.02×10^{23} atoms/ 1 mole) = 4.42×10^{23} atoms Zn. 3. 452 g Ar \times (1 mol Ar/ 39.95g) = 11.3 moles Ar. 4. 0.05 mol \times (22.4 L / 1 mol) = 1.12 L.

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carbon-12 (which is Avogadro's number, or 6.022×10^{23}).

Avagadro's number. the constant, 6.022×10^{23} , representing the number of atoms, molecules, or ions in one mole of a substance. Moles Study Guide Chemistry Flashcards | Quizlet
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(which is Avogadro's number, or Page 7/24

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Chapter 10 Chemical Calculations and Chemical Equations

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Guide The mole is the most common
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chemical substance For all solids,

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liquids, and gases, you can convert mass to moles (or moles to mass) For gases, you should memorize the following conversion of volume to moles (or moles to

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Toulouse III (France) and a PhD in Chemical Engineering at the University of South Carolina (USA) in 2008. Dr. Jerome Delhommelle earned his PhD in Chemistry at the University of Paris XI-Orsay (France) in 2000. He is currently working as an Associate Professor in Chemistry at the

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Study Guide to Accompany Basics for

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are a recently discovered allotrope of carbon. Three other allotropes of carbon-buckyballs, graphite, and diamond-are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds

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